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Engineering Software is pleased to announce the introduction of *Free Coursework Material*.

Engineering Software Coursework Material covers the following area:

 *Power Cycle Components/Processes*
Compression
Stoichiometric Combustion (Carbon, Hydrogen, Sulfur, Coal, Oil and Gas)
Expansion

Compression

Introduction

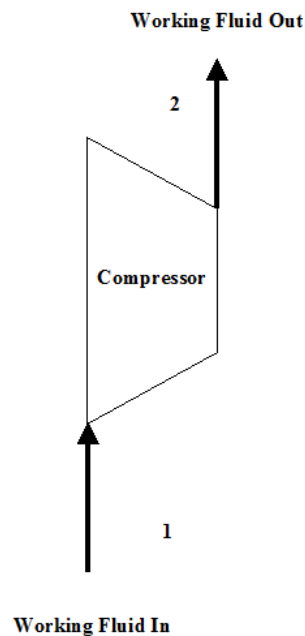
This section provides an isentropic compression analysis when the working fluid is air.

Analysis

In the presented compression analysis, only air is considered as the working fluid behaving as a perfect gas -- specific heat has a constant value. Ideal gas state equation is valid -- $pv = RT$.

Air enters a compressor at point 1 and it exits the compressor at point 2. Isentropic compression is considered with no entropy change.

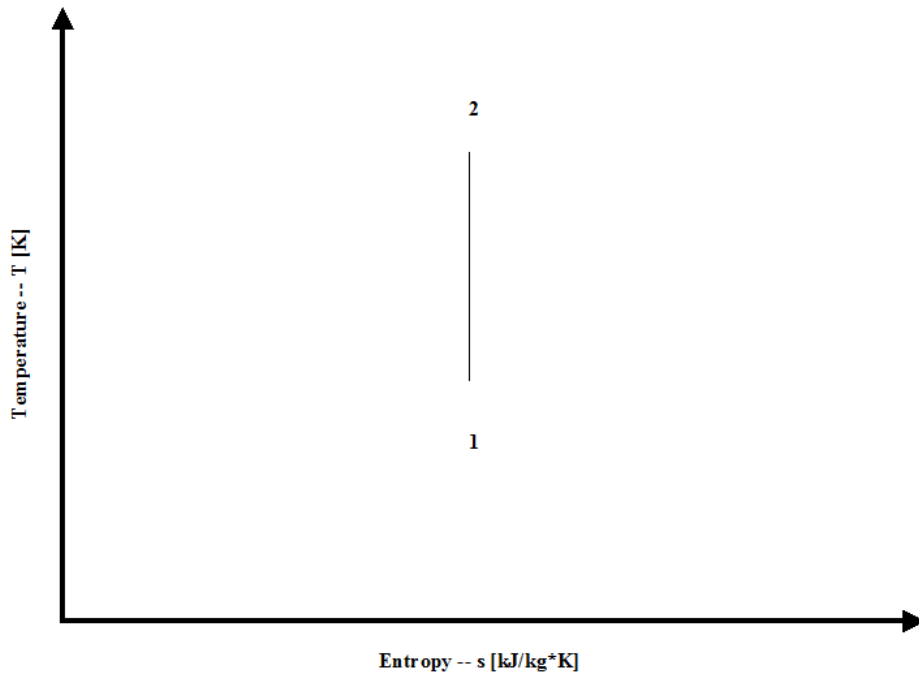
Figure 1 presents a compression schematic layout.



Compression Schematic Layout

Figure 1 - Compression Schematic Layout

Figure 2 presents a compression temperature vs entropy diagram.



Compression T - s Diagram

Figure 2- Compression Temperature vs Entropy Diagram

Figure 3 presents compression specific power input requirements for a few typical compression ratio values. It should be noted that the air enters the compressor at standard ambient conditions of 298 [K] and 1 [atm] of absolute pressure.

Compression Specific Power Input

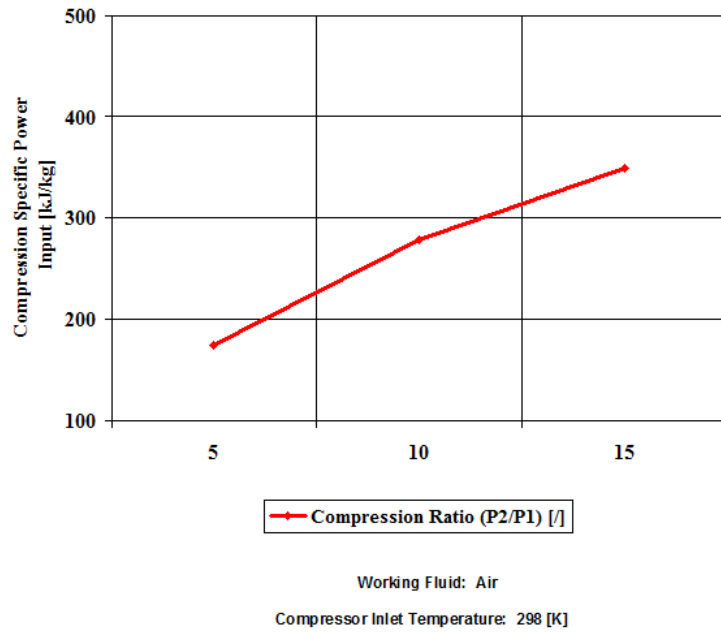


Figure 3 - Compression Specific Power Input

Figure 4 presents compression power input requirements for two typical compression ratio values and a few different working fluid mass flow rate values.

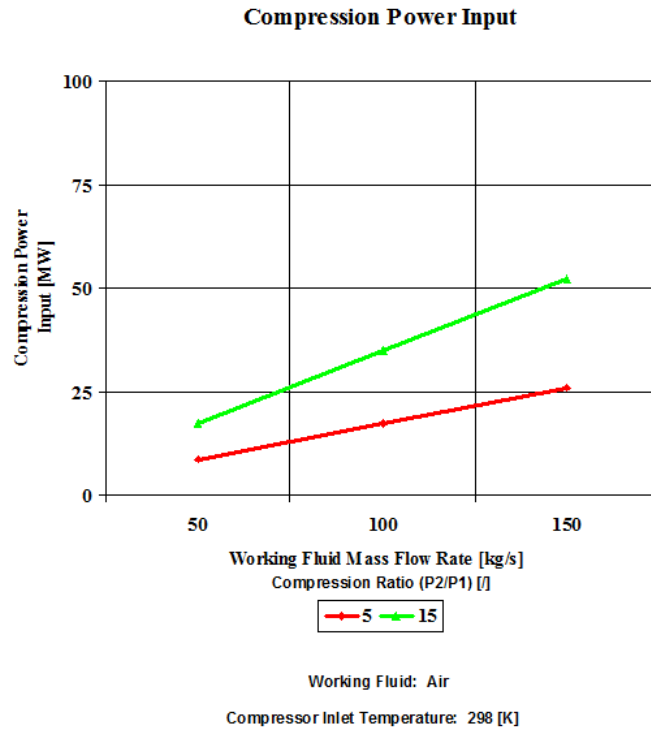


Figure 4 - Compression Power Input

One can notice that both compression specific power input and power input increase with the compression ratio. As the working fluid mass flow rate increases, the compression power input requirements increase too.

Assumptions

Working fluid is air. There is no friction and heat transfer. Compression is isentropic -- there is no entropy change. Ideal gas state equation is valid -- $pv = RT$. Air behaves as a perfect gas -- specific heat has a constant value.

Governing Equations

$$T_2/T_1 = (p_2/p_1)^{(\kappa-1)/\kappa}$$

$$\kappa = c_p/c_v$$

$$pv = RT$$

$$w = c_p(T_2 - T_1)$$

$$W = c_p(T_2 - T_1)m$$

Input Data

$$T_1 = 298 \text{ [K]}$$

$$p_1 = 1 \text{ [atm]}$$

$$p_2 = 5, 10 \text{ and } 15 \text{ [atm]}$$

$$c_p = 1.004 \text{ [kJ/kg}\cdot\text{K]}$$

$$\kappa = c_p/c_v - \text{for air } \kappa = 1.4 \text{ []}$$

$$m = 50, 100 \text{ and } 150 \text{ [kg/s]}$$

Results

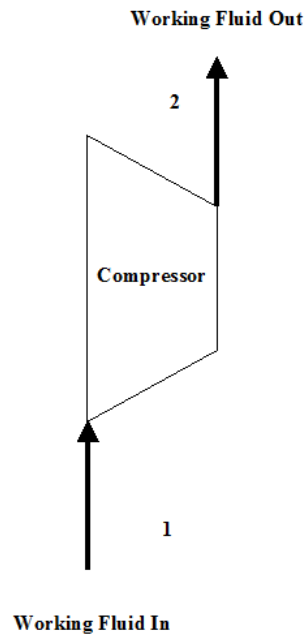
Specific Power Input vs Compression Ratio

Compression Ratio [/]	Specific Power Input [kW/kg/s]
5	174
10	278
15	349

Power Input vs Compression Ratio for a few Mass Flow Rates

Power Input [MW]	Mass Flow Rate [kg/s]		
Compression Ratio [/]	50	100	150
5	8.7	17.4	26.0
15	17.5	34.9	52.4

Figures

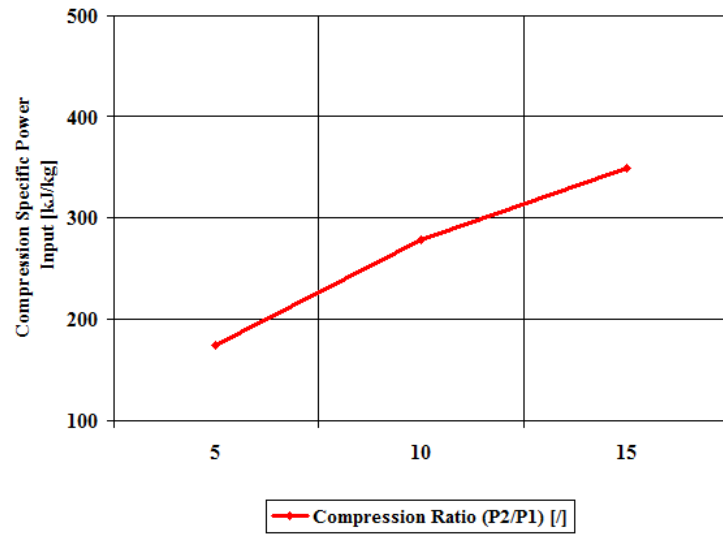


Compression Schematic Layout

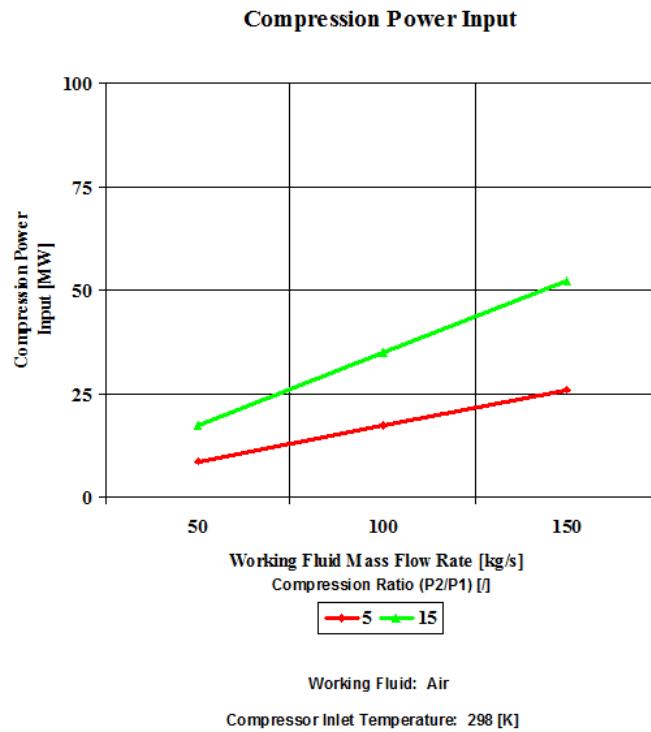


Compression T - s Diagram

Compression Specific Power Input



Working Fluid: Air
Compressor Inlet Temperature: 298 [K]



Conclusions

Both compression specific power input and power input increase with the compression ratio. As the working fluid mass flow rate increases, the compression power input requirements increase too.

References

JANAF Thermochemical Data - Tables, 1970

Stoichiometric Combustion (Carbon, Hydrogen, Sulfur, Coal, Oil and Gas)

Introduction

This section provides a combustion analysis for a few typical fuel cases (carbon, hydrogen, sulfur, coal, oil and gas) when the fuel reacts with air at stoichiometric conditions.

Analysis

In the presented combustion analysis, both fuel and air are at standard inlet combustion conditions of 298 [K] and 1 [atm] of absolute pressure. Furthermore, combustion is complete and with no heat loss.

During combustion, a large amount of reactants' chemical energy gets released in the form of thermal energy.

Fuel higher heating value (HHV) or heat of combustion is the difference between the reactants enthalpy value minus the combustion products enthalpy value at the standard reference temperature, which is 298 [K].

When the reactants enthalpy value is equal to the combustion products enthalpy value, one can calculate the combustion products flame temperature or adiabatic temperature.

Figure 1 presents the reactants and combustion products enthalpy value change with an increase in the temperature.

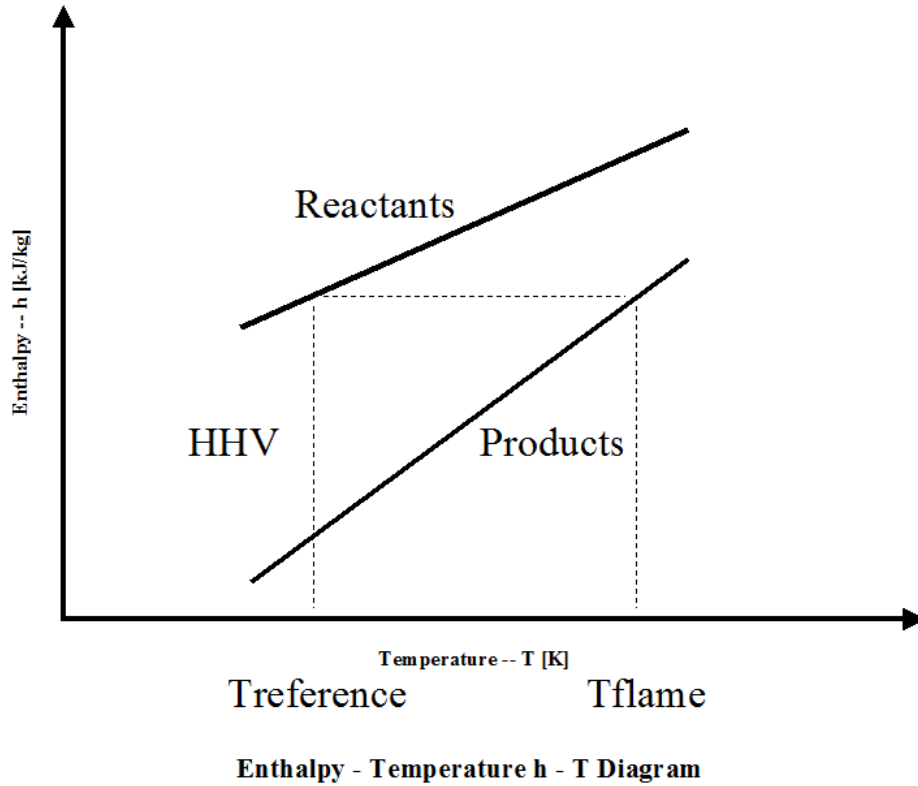


Figure 1 - Reactants and Combustion Products Enthalpy vs Temperature

Physical properties for both reactants and combustion products are very important and need to be known in order to carry out successful combustion calculations.

Figure 2 presents how the reactants and combustion products species enthalpy values change with the temperature. The physical properties provided in Figure 2 come from the JANAF Thermochemical Data - Tables, 1970.

Enthalpy vs Temperature

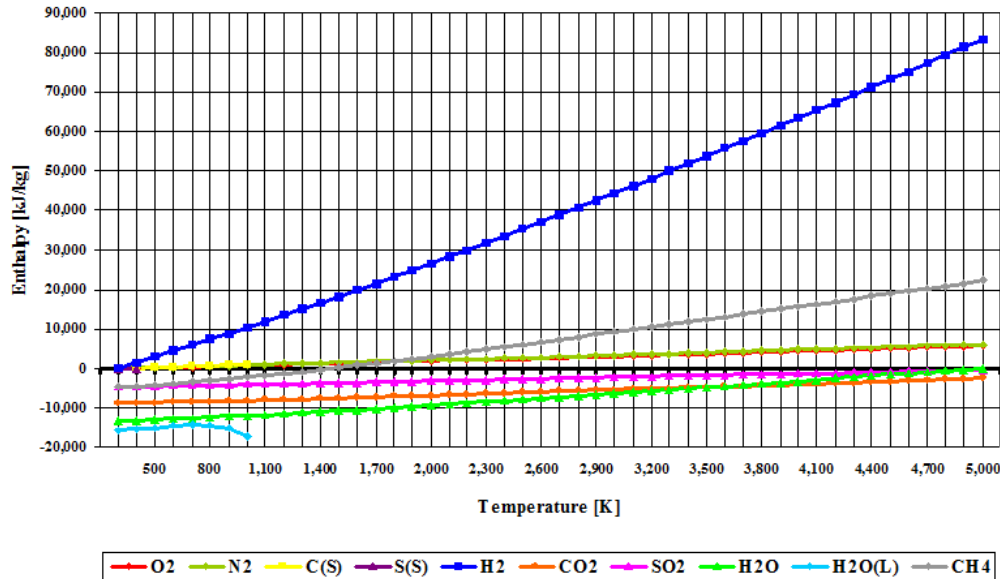


Figure 2 - Reactants and Combustion Products Species Enthalpy vs Temperature

It is interesting to note that the enthalpy value for basic combustion elements such as carbon (C), hydrogen (H₂), sulfur (S), oxygen (O₂) and nitrogen (N₂) is equal to zero at the standard combustion conditions of 298 [K] and 1 [atm].

Also, it should be mentioned that for ideal gas species, the enthalpy value is only dependent on the temperature.

In addition to knowing the reactants and combustion products physical properties, for any kind of combustion analysis and calculations, it is important to know both fuel and oxidant compositions.

For solid and liquid type fuels, the fuel composition is given on the weight basis for a unit mass amount. For the gas type fuels, the fuel composition is provided on the mole/volume basis for a unit volume amount. In this analysis, methane (CH₄) is the only gas fuel considered. In order to keep the combustion analysis simple and straightforward, the CH₄ composition is provided on the weight basis. Oxidant composition is usually given on the mole/volume basis.

Table 1 provides the fuel composition.

Table 1 - Fuel Composition

Fuel	C [kg/kg]	H [kg/kg]	S [kg/kg]	N [kg/kg]	O [kg/kg]	H₂O [kg/kg]	CH₄ [kg/kg]
Carbon	1.000	0.000	0.000	0.000	0.000	0.000	-
Hydrogen	0.000	1.000	0.000	0.000	0.000	0.000	-
Sulfur	0.000	0.000	1.000	0.000	0.000	0.000	-
Coal	0.780	0.050	0.030	0.040	0.080	0.020	-
Oil	0.860	0.140	0.000	0.000	0.000	0.000	-
Gas	-	-	-	-	-	-	1.000

Table 2 provides the standard air composition.

Table 2 - Oxidant Composition

Oxidant	N [kg/kg]	O [kg/kg]	N [kmol/kmol]	O [kmol/kmol]
Air	0.767	0.233	0.790	0.210

Again, in this combustion analysis, only the stoichiometric combustion is analyzed. Results of such analysis are provided, including combustion products composition on weight and mole/volume basis, flame temperature, stoichiometric oxidant to fuel ratio and fuel higher heating value (HHV).

Table 3 provides the combustion products composition on the weight basis.

Table 3 - Combustion Products Composition on the Weight Basis

Fuel	CO₂ [kg/kg]	H₂O [kg/kg]	SO₂ [kg/kg]	N₂ [kg/kg]	O₂ [kg/kg]
Carbon	0.295	0.000	0.000	0.705	0.000
Hydrogen	0.000	0.255	0.000	0.745	0.000
Sulfur	0.000	0.000	0.378	0.622	0.000
Coal	0.249	0.041	0.005	0.705	0.000
Oil	0.202	0.080	0.000	0.718	0.000
Gas	0.151	0.124	0.000	0.725	0.000

Table 4 provides the combustion products composition on the volume basis.

Table 4 - Combustion Products Composition on the Volume Basis

Fuel	CO₂ [kmol/kmol]	H₂O [kmol/kmol]	SO₂ [kmol/kmol]	N₂ [kmol/kmol]	O₂ [kmol/kmol]
Carbon	0.210	0.000	0.000	0.790	0.000
Hydrogen	0.000	0.347	0.000	0.653	0.000
Sulfur	0.000	0.000	0.210	0.790	0.000
Coal	0.170	0.068	0.002	0.759	0.000
Oil	0.132	0.129	0.000	0.739	0.000
Gas	0.095	0.190	0.000	0.715	0.000

When considering coal, oil and gas as the fuel, coal has the largest amount of CO₂ in the combustion products on both weight and mole basis.

Table 5 provides combustion products flame temperature, stoichiometric oxidant to fuel ratio and fuel higher heating value.

Table 5 - Combustion Products Flame Temperature, Stoichiometric Oxidant to Fuel Ratio and Fuel Higher Heating Value

Fuel	Flame Temperature [K]	Stoichiometric Oxidant to Fuel Ratio [/]	HHV [Btu/lbm]
Carbon	2,460	11.444	14,094
Hydrogen	2,525	34.333	60,997
Sulfur	1,972	4.292	3,982
Coal	2,484	10.487	14,162
Oil	2,484	14.649	20,660
Gas	2,327	17.167	21,563

Stoichiometric oxidant to fuel ratio is the mass of air required for complete combustion of a unit mass of fuel. Thus, 1 [kg] of carbon fuel requires 11.433 [kg] of air for complete, ideal combustion.

Today, global warming is becoming more evident and it is being said that it is primarily caused by CO₂ emissions. A detailed combustion analysis, as it is provided here, can be very useful in

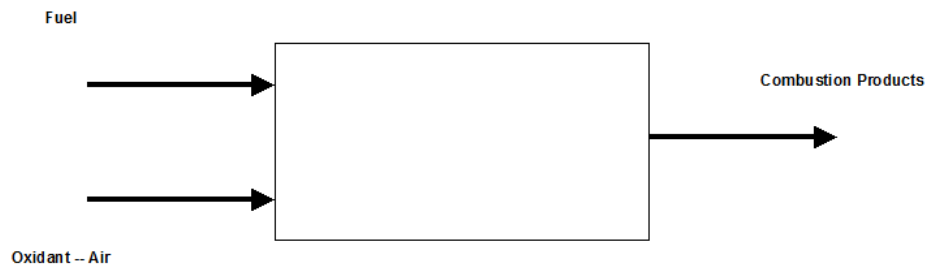
determining different fuel and technology scenarios that would result in the reduction of current CO₂ emissions.

Assumptions

Both fuel and air are at standard inlet combustion conditions of 298 [K] and 1 [atm] of absolute pressure.

Furthermore, combustion is complete and with no heat loss.

Combustion Schematic Layout



Species Molecular Weight

Species	C	H ₂	S	O ₂	N ₂	CO ₂	H ₂ O	SO ₂	CH ₄
Molecular Weight [kg/kmol]	12	2	32	32	28	44	18	64	16

**Stoichiometric Fuel, Oxidant and Combustion Products
Volume Flow Rates for 1 [kmol] of Fuel Species**

Fuel	Fuel Composition				Oxidant Composition		Combustion Products Composition				
	C [kmol]	H [kmol]	S [kmol]	CH ₄ [kmol]	O ₂ [kmol]	N ₂ [kmol]	CO ₂ [kmol]	H ₂ O [kmol]	SO ₂ [kmol]	O ₂ [kmol]	N ₂ [kmol]
Carbon	1	0	0	0	1	3.762	1	0	0	0	3.762
Hydrogen	0	1	0	0	0.5	1.881	0	1	0	0	1.881
Sulfur	0	0	1	0	1	3.762	0	0	1	0	3.762
Coal	1	1	1	0	2.5	9.404	1	1	1	0	9.404
Oil	1	1	0	0	1.5	5.642	1	1	0	0	5.642
Gas	0	0	0	1	2	7.524	1	2	0	0	7.524

**Stoichiometric Fuel, Oxidant and Combustion Products
Mass Flow Rates for 1 [kmol] of Fuel Species**

Fuel	Fuel Composition				Oxidant Composition		Combustion Products Composition					Stoichiometric Ratio [l]
	C [kg]	H [kg]	S [kg]	CH ₄ [kg]	O ₂ [kg]	N ₂ [kg]	CO ₂ [kg]	H ₂ O [kg]	SO ₂ [kg]	O ₂ [kg]	N ₂ [kg]	
Carbon	12	0	0	0	32	105.336	44	0	0	0	105.336	11.444
Hydrogen	0	2	0	0	16	52.668	0	18	0	0	52.668	34.333
Sulfur	0	0	32	0	32	105.336	0	0	64	0	105.336	4.292
Coal	12	2	32	0	80	263.312	44	18	64	0	263.312	7.463
Oil	12	2	0	0	48	157.976	44	18	0	0	157.976	14.712
Gas	0	0	0	16	64	210.672	44	36	0	0	210.672	17.167

Governing Equations

Fuel higher heating value (HHV) or heat of combustion is the difference between the reactants enthalpy value minus the combustion products enthalpy value at the standard reference temperature, which is 298 [K].

When the reactants enthalpy value is equal to the combustion products enthalpy value, one can calculate the combustion products flame temperature or adiabatic temperature.

Input Data

Table 1 - Fuel Composition

Fuel	C [kg/kg]	H [kg/kg]	S [kg/kg]	N [kg/kg]	O [kg/kg]	H₂O [kg/kg]	CH₄ [kg/kg]
Carbon	1.000	0.000	0.000	0.000	0.000	0.000	-
Hydrogen	0.000	1.000	0.000	0.000	0.000	0.000	-
Sulfur	0.000	0.000	1.000	0.000	0.000	0.000	-
Coal	0.780	0.050	0.030	0.040	0.080	0.020	-
Oil	0.860	0.140	0.000	0.000	0.000	0.000	-
Gas	-	-	-	-	-	-	1.000

Table 2 - Oxidant Composition

Oxidant	N [kg/kg]	O [kg/kg]	N [kmol/kmol]	O [kmol/kmol]
Air	0.767	0.233	0.790	0.210

Results

Table 3 - Combustion Products Composition on the Weight Basis

Fuel	CO₂ [kg/kg]	H₂O [kg/kg]	SO₂ [kg/kg]	N₂ [kg/kg]	O₂ [kg/kg]
Carbon	0.295	0.000	0.000	0.705	0.000
Hydrogen	0.000	0.255	0.000	0.745	0.000
Sulfur	0.000	0.000	0.378	0.622	0.000
Coal	0.249	0.041	0.005	0.705	0.000
Oil	0.202	0.080	0.000	0.718	0.000
Gas	0.151	0.124	0.000	0.725	0.000

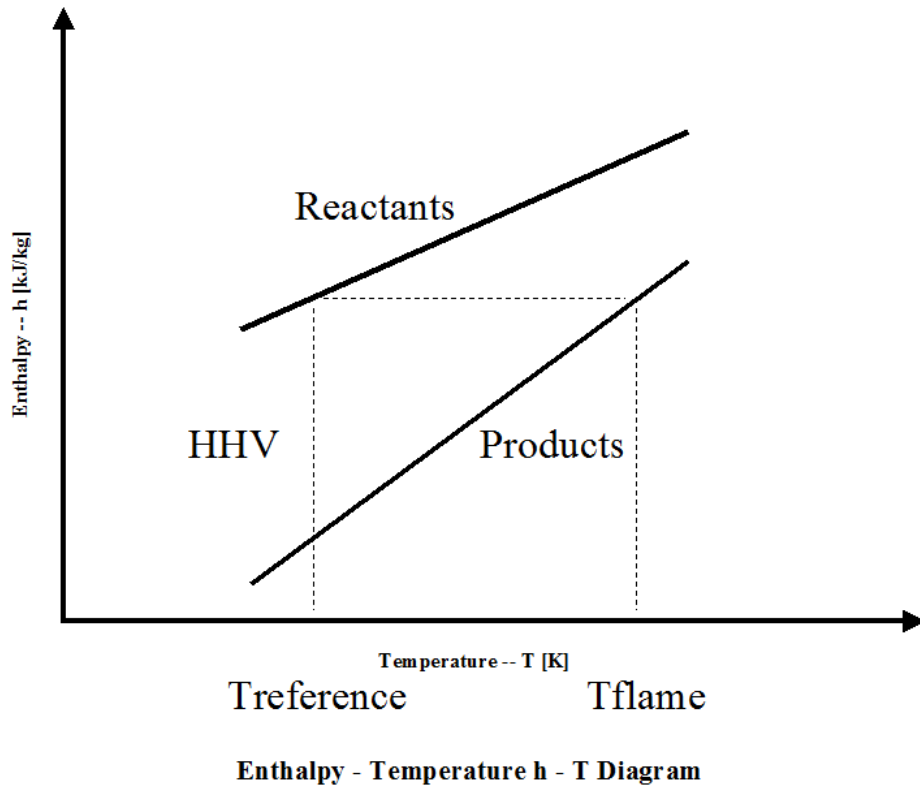
Table 4 - Combustion Products Composition on the Volume Basis

Fuel	CO₂ [kmol/kmol]	H₂O [kmol/kmol]	SO₂ [kmol/kmol]	N₂ [kmol/kmol]	O₂ [kmol/kmol]
Carbon	0.210	0.000	0.000	0.790	0.000
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Sulfur	0.000	0.000	0.210	0.790	0.000
Coal	0.170	0.068	0.002	0.759	0.000
Oil	0.132	0.129	0.000	0.739	0.000
Gas	0.095	0.190	0.000	0.715	0.000

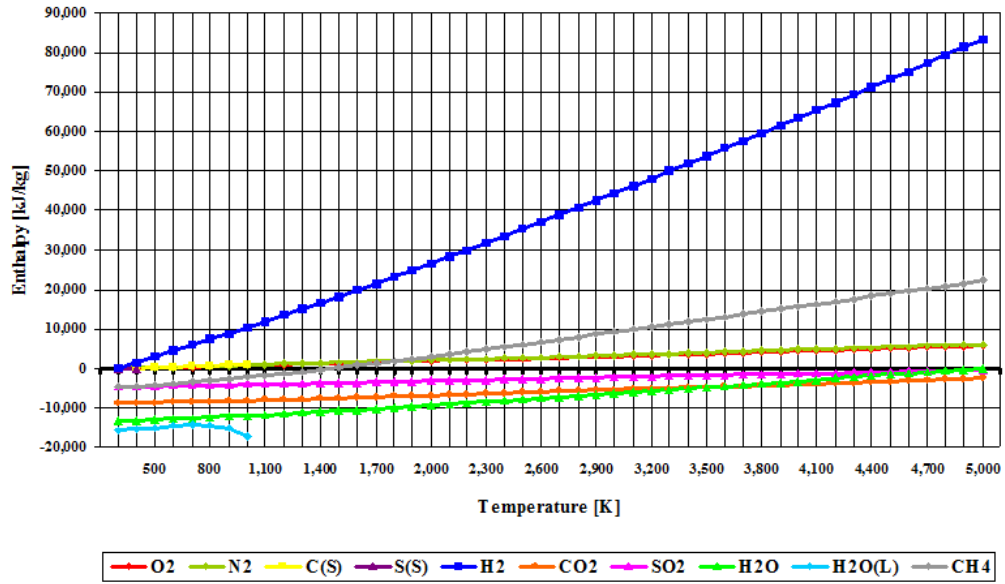
Table 5 - Combustion Products Flame Temperature, Stoichiometric Ratio and the Fuel Higher Heating Value

Fuel	Flame Temperature [K]	Stoichiometric Ratio [/]	HHV [Btu/lbm]
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Hydrogen	2,525	34.333	60,997
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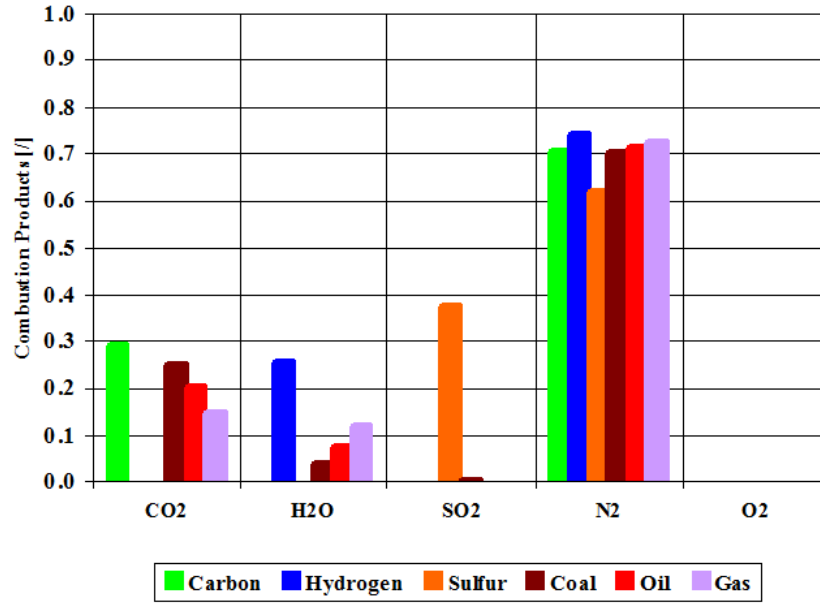
Figures



Enthalpy vs Temperature

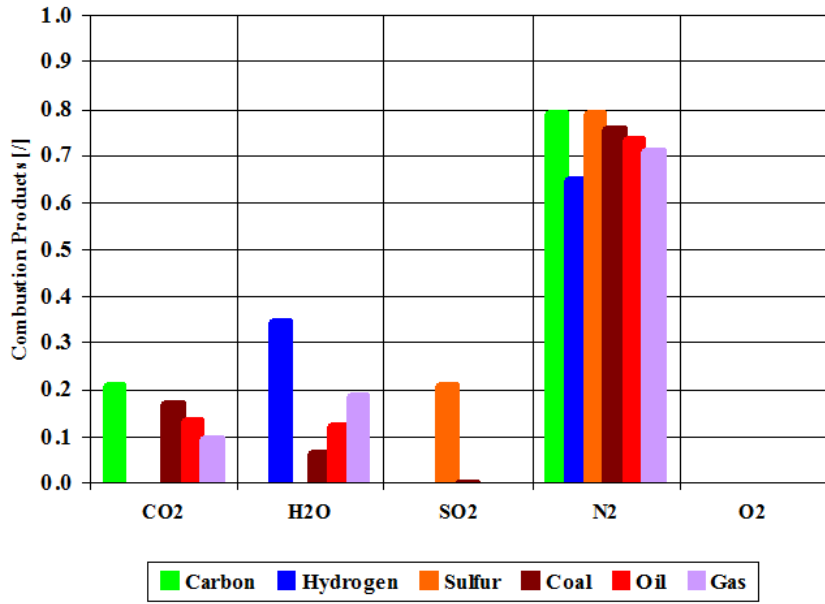


Combustion Products -- Weight Basis



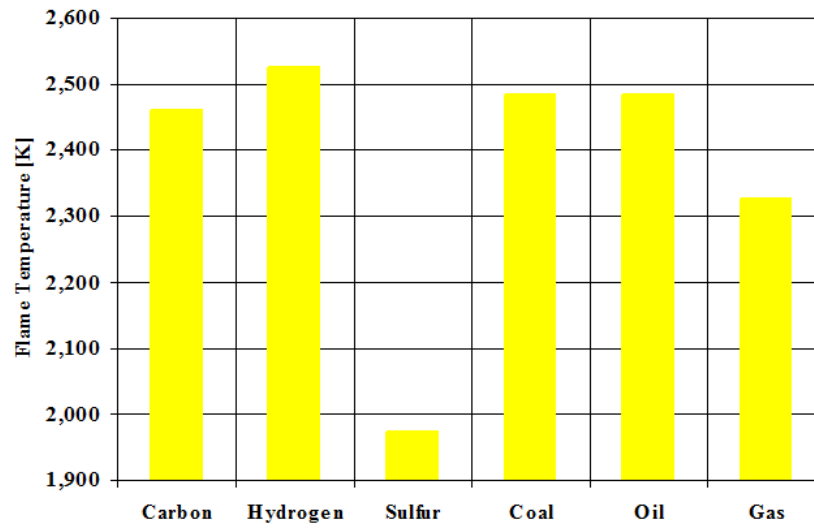
Fuel and Oxidant Inlet Temperature: 298 [K]

Combustion Products -- Mole Basis



Fuel and Oxidant Inlet Temperature: 298 [K]

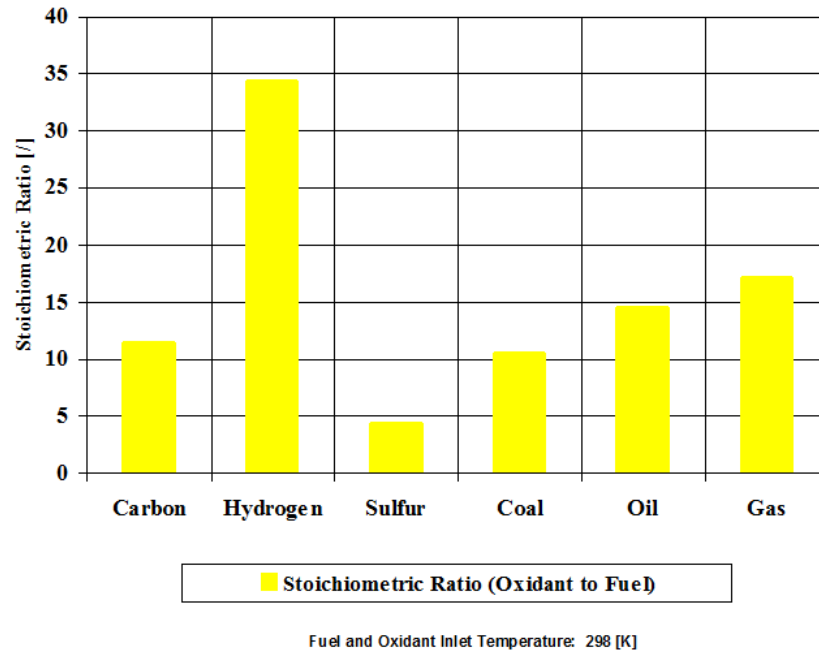
Combustion Products Flame Temperature



Flame Temperature

Fuel and Oxidant Inlet Temperature: 298 [K]

Combustion Stoichiometric Ratio



Conclusions

Hydrogen as the fuel has the highest flame temperature, requires the most mass amount of oxidant/air in order to have complete combustion per unit mass amount of fuel and the largest higher heating fuel value.

When hydrogen reacts with oxidant/air, there is no CO₂ present in the combustion products.

References

JANAF Thermochemical Data - Tables, 1970

Expansion

Introduction

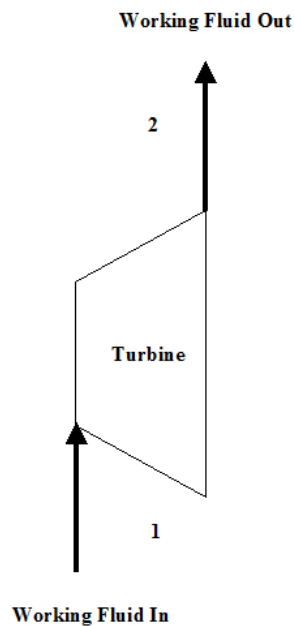
This section provides an isentropic expansion analysis when the working fluid is air.

Analysis

In the presented expansion analysis, only air is considered as the working fluid behaving as a perfect gas -- specific heat has a constant value. Ideal gas state equation is valid -- $pv = RT$.

Air enters a turbine at point 1 and it exits the turbine at point 2. Isentropic expansion is considered with no entropy change.

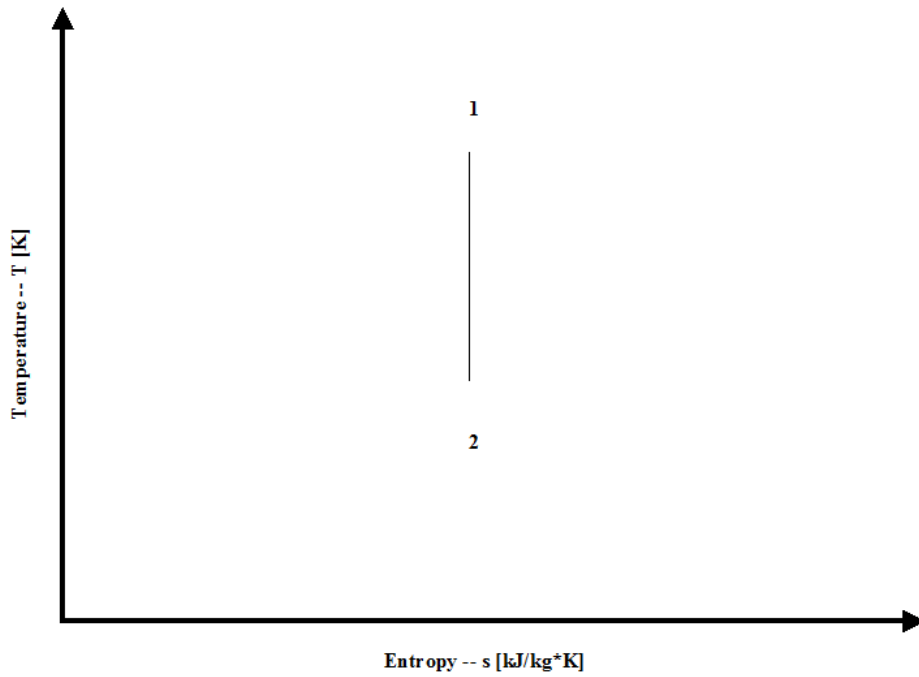
Figure 1 presents an expansion schematic layout.



Expansion Schematic Layout

Figure 1 - Expansion Schematic Layout

Figure 2 presents an expansion temperature vs entropy diagram.



Expansion T - s Diagram

Figure 2 - Expansion Temperature vs Entropy Diagram

Figure 3 presents expansion specific power output values for a few typical expansion ratio values. It should be noted that the air enters the turbine at the temperature of 1,500 [K] and the turbine exhaust pressure is always equal to the standard ambient pressure -- 1 [atm] of absolute pressure.

Expansion Specific Power Output

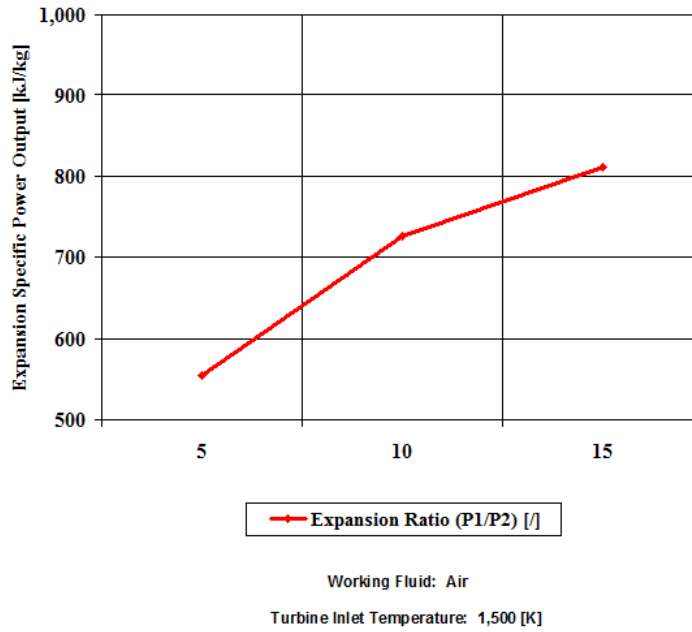


Figure 3 - Expansion Specific Power Output

Figure 4 presents expansion power output values for two typical expansion ratio values and a few different working fluid mass flow rate values.

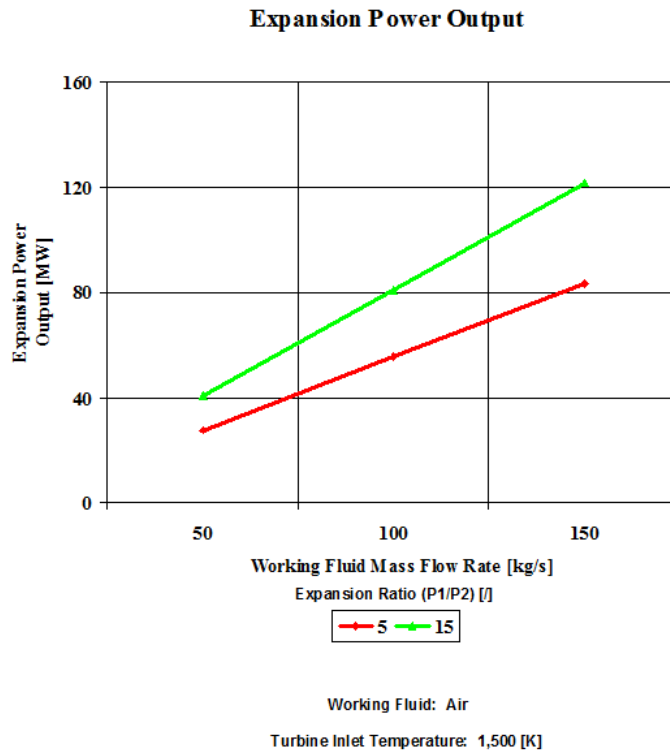


Figure 4 - Expansion Power Output

One can notice that both expansion specific power output and power output increase with the expansion ratio. As the working fluid mass flow rate increases, the expansion power output values increase too.

Assumptions

Working fluid is air. There is no friction and heat transfer. Expansion is isentropic -- there is no entropy change. Ideal gas state equation is valid -- $p v = R T$. Air behaves as a perfect gas -- specific heat has a constant value.

Governing Equations

$$T_1/T_2 = (p_1/p_2)^{(\kappa-1)/\kappa}$$

$$\kappa = c_p/c_v$$

$$pv = RT$$

$$w = c_p(T_1 - T_2)$$

$$W = c_p(T_1 - T_2)m$$

Input Data

$$T_1 = 298 \text{ [K]}$$

$$p_1 = 5, 10 \text{ and } 15 \text{ [atm]}$$

$$p_2 = 1 \text{ [atm]}$$

$$c_p = 1.004 \text{ [kJ/kg}\cdot\text{K]}$$

$$\kappa = c_p/c_v - \text{for air } \kappa = 1.4 \text{ [/]}$$

$$m = 50, 100 \text{ and } 150 \text{ [kg/s]}$$

Results

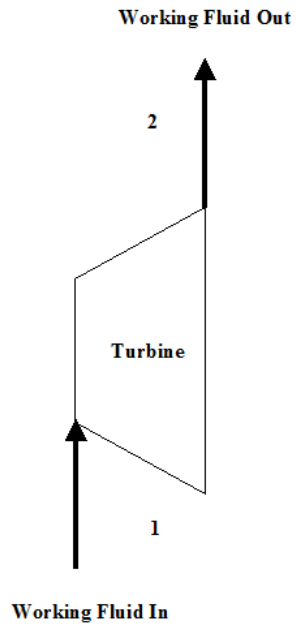
Specific Power Output vs Expansion Ratio

Expansion Ratio [/]	Specific Power Output [kW/kg/s]
5	555
10	726
15	811

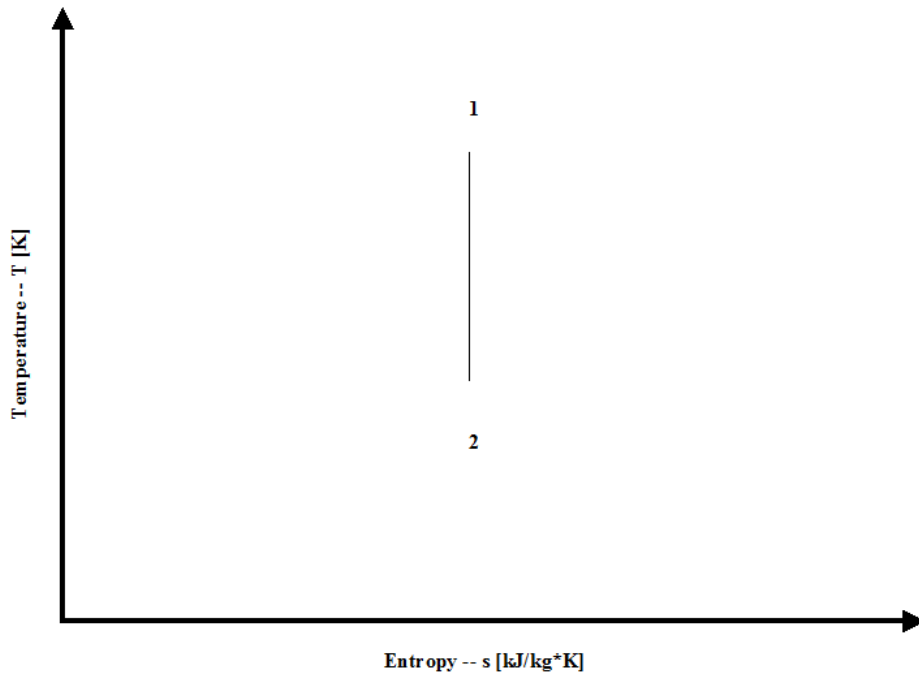
Power Output vs Expansion Ratio for a few Mass Flow Rates

Power Output [MW]	Mass Flow Rate [kg/s]		
Expansion Ratio [/]	50	100	150
5	27.8	55.5	83.3
15	40.6	81.1	121.7

Figures

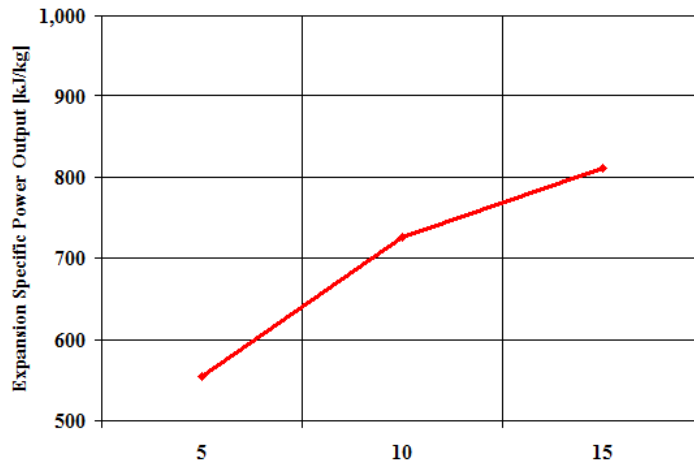


Expansion Schematic Layout



Expansion T - s Diagram

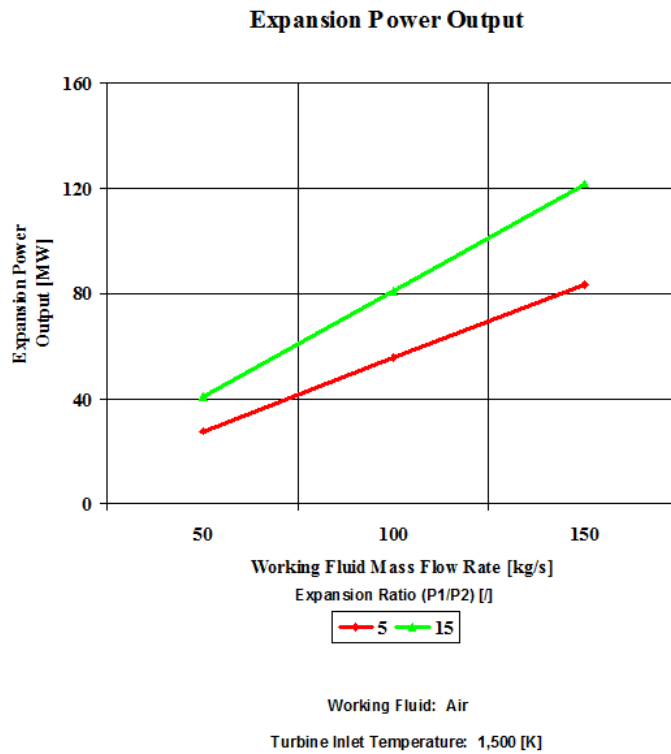
Expansion Specific Power Output



Expansion Ratio (P1/P2) []

Working Fluid: Air

Turbine Inlet Temperature: 1,500 [K]



Conclusions

Both expansion specific power output and power output increase with the expansion ratio. As the working fluid mass flow rate increases, the expansion power output values increase too.

References

JANAF Thermochemical Data - Tables, 1970